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Smectite quantification in hydrothermally altered volcanic rocks

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A B S T R A C T

In volcanic environments, the presence of smectite may indicate recent hydrothermal circulations. Smectite is also responsible for enhanced rock electrical conductivity, as well as mechanical weakening. Therefore, quantifying smectite is important in geothermal exploration. Smectite identification requires X-ray diffraction (XRD) but quantification based on XRD is time-consuming and not always accurate. In the present study, we investigate the use of an optimized unbuffered Cation Exchange Capacity (CEC) determination, by back-titration of the Copper-triethylenetetramine(II) "Cu-trien" molecule, to quantify the smectite content of altered volcanic rock samples. We establish that a satisfying trade-off between the instrument uncertainty and an independant systematic error is theoretically reached for a fraction of reactants consumed of about 30% at the end of the exchange reaction. We suggest a modification to classical protocols to fall in that range. Finally, we show that optimized CEC determination by Cu-trien are a direct measure of the smectite weight fraction in altered volcanic samples, with an average CEC of pure smectite of 90 ± 5 meq/100 g.

1. Introduction

One of the challenges of geothermal exploration at volcanoes is to detect the presence of active hydrothermal circulations in fractures. Geo-electrical and electromagnetic measurements are commonly used to this aim because electrical resistivity contrasts can delineate zones of intense hydrothermal activity (e.g. Árnason et al., 2000; Flóvenz et al., 2005; Flóvenz et al., 2012). Electrical resistivity of volcanic rocks is particularly sensitive to the presence of secondary "alteration" minerals, often evidences of hydrothermal circulations in fractures, such as clay minerals. The distribution of clay minerals can provide estimates of the temperature distribution in volcanic or sedimentary systems where their formation is controlled by the geothermal gradient (Alt et al., 1986; Bourdelle et al., 2013; Kristmannsdóttir and Tómasson, 1978; Kristmannsdottir, 1979). In active hydrothermal systems, the formation of smectite is not only controlled by the geothermal gradient, but also by the convective activity related to recent fault opening and causing boiling as well as chemical disequilibrium (Beaufort et al., 1995; Bril et al., 1996; Patrier et al., 1996). Compared to other clay minerals (e.g. illite, chlorite, kaolinite), smectite is much more conductive (e.g. Kaufhold et al., 2014; Kaufhold et al., 2015) and contributes significantly to the electrical conductivity of rocks, through Electrical Double Layer mechanisms (Flóvenz et al., 1985; Pezard, 1990; Revil and Glover, 1997; Waxman and Smits, 1968) and interfoliar conduction (Henry, 1997; Lévy et al., 2018; Maraqah et al., 1990). Smectite is also abundant in subduction zones (Hyndman et al., 1997) and in some major faults (Chester et al., 2013) and may play a role in the mechanical weakening of altered volcanic rocks (Heap et al., 2014; Kaufhold et al., 2012; Meller, 2014).

In order to study in the laboratory the influence of smectite on electrical conductivity and mechanical properties of volcanic rocks, the smectite content needs to be quantified first. Smectite content in drill-cuttings can also provide estimates of the porosity or permeability in a reservoir, by comparison with in-situ borehole resistivity logs (Flóvenz et al., 2005; Pezard, 1990; Revil et al., 1998; Rink and Schopper, 1974; Waxman and Smits, 1968).

The primary goal of our study is to provide geothermal industry with a simple method to quantify smectite content in hydrothermally altered volcanic rocks. Quantifying smectite in altered volcanic rocks is challenging because a large number of minerals often coexists in the same rock formation. Quantification of minerals by Rietveld-refinements of X-ray diffraction (XRD) patterns (e.g. Taut et al., 1998) can generally be applied to altered volcanic rocks (Lévy et al., 2018). But smectite quantification is hampered when smectite-containing mixed layers co-exist with smectite (Raven and Self, 2017). Moreover, high-quality XRD scans are required for these quantifications, which can be

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time-consuming.

Due to its particular crystalline structure, smectite has a much larger Cation Exchange Capacity (CEC) than other clay minerals (Bouchet et al., 2000). This CEC is mainly located in smectite interlayers (Dohrmann, 2006a, 2006b; Lagaly, 1981; Vogt and Köster, 1978). Hower and Mowatt (1966) found a linear correlation between the CEC and the smectite fraction in a series of illite-smectite samples. Kaufhold and Dohrmann (2003) also observed that the CEC determined by back-titration of the Copper-triethylenetetramine(II) "Cu-trien" (Ammann et al., 2005; Bergaya and Vayer, 1997; Meier and Kahr, 1999) was proportional to the smectite content in bentonites, qualitatively determined using the methylene blue method.

Altered volcanic rocks contain a larger variety of minerals than bentonite or illite-smectite series, with often a large fraction of zeolites. Some zeolites, such as clinoptilolite and heulandite, have a higher CEC than smectite, up to 300 meq/100 g, thanks to their wide solid solution of extraframework cations (Fridriksson et al., 2004). However, the CEC of clinoptilolite drops to 5 meq/100 g when determined by the Cu-trien method (Meier and Kahr, 1999), because the channels where extra framework cations are located cannot expand, unlike smectite interlayers, so that only small cations (smaller than Cu-trien) can enter clinoptilolite channels. Therefore, the Cu-trien molecule appears to be adequate to quantify the smectite content in altered volcanic rocks. Our study tests this possibility by comparing CEC determinations to smectite quantifications based on Rietveld-refinements of XRD patterns, for samples where smectite is the only swelling clay mineral.

Since altered volcanic rocks contain a lower and more variable smectite content than bentonite or illite-smectite series, the solid/solution ratio needs to be optimized for each sample, in order to minimize both the instrument uncertainty and systematic biases. The need for optimization of the reactants (exchange solution and sample) proportion was first addressed by Orsini and Remy (1976) for CEC determinations with the Cobalti-hexamine molecule, "Co-hex", on large masses of soil samples. Orsini and Remy (1976) pointed out that the exchange between Co-hex and soil samples could be considered as total only if the initial quantity of Co-hex was at least three times greater than the number of exchange sites (the CEC in appropriate units). Yet, beyond eight times, the accuracy of the determinations would significantly decrease. Based on these observations, these authors recommended carrying out experiments where the initial ratio between Co-hex and rock sample (expressed in CEC units) represents 30% to 80% of the CEC, or equivalently where the fraction of Co-hex consumed during the experiment ranges between 15% and 30%. Further development of the Co-hex back-titration method (Ciesielski et al., 1997a, 1997b) allowed extending the interval of Co-hex consumed to 5%–35%. The need to optimize the solid/solution ratio by adjusting the initial mass of sample, for samples having a wide range of CEC values, is also discussed for the Cu-trien method by Dohrmann and Kaufhold (2009) and by Dohrmann (2006) for the similar Ag-thiourea method.

Our study investigates the theoretical grounds for the observations by Dohrmann and Kaufhold (2009) and the ranges suggested by Orsini and Remy (1976) and Ciesielski et al. (1997) and presents a simple method for quantifying the smectite content of altered volcanic rocks through optimized CEC determinations using the Cu-trien exchange complex.

2. Materials and methods

2.1. Rock samples

Thirty-eight samples from the Krafla high-temperature geothermal area are used in this study. Core samples are collected from four cored boreholes (KH1, KH3, KH5 and KH6) at varying depths. They represent a variety of lithologies and secondary minerals (Table 1). Cylindrical plugs (2–3 cm long and 2.5 cm diameter) are prepared from the original core samples for petrophysical measurements, described by Lévy et al. (2018). From the lateral faces of the plugs, thin sections and powders are prepared. Ten samples are used for optimization of CEC determination, 24 samples for comparison between quantitative XRD analysis and CEC, 15 samples for chemical analysis by Electron Probe Micro Analysis (EPMA) and four samples for ICP analysis of exchangeable cations. Some samples are used for more than one type of analysis (Table 1).

A first set of powders is used for the analysis of CEC uncertainty, carried out on 88 samples but presented here for 10 relevant samples. This set of powders is roughly grained to powder size, without any size control.

A second set of powders is used for XRD scans and associated mineral quantification by Rietveld-refinements. Powders from the first set are further ground for 10 minutes in ethanol, using an automatic grinder Retsch RM 200. These finer powders are dried, sieved and then prepared as randomly oriented mounts. For each sample, the exact same powder is used later on for independent CEC determinations that are compared to XRD mineral quantification.

2.2. Mineral quantification by X-ray diffraction

The powders are front-loaded onto the sample holder, using a razor blade to smooth out the surface, in order to minimize the preferred orientation (PO) and shift of diffraction peaks (Bish and Reynolds, 1989). An example of PO issues when the sample is back-loaded is shown in Appendix A. The XRD scans are carried out over the range 4–75°2θ (4-65°2θ for a few samples) with a Philips X'Pert Pro (radiation Cu-Kα; 45 kV; 30 mA; step size 0.0167°; time per step 240 s; X'Celerator Scientific high-speed detector; 240 mm goniometer radius). The XRD patterns are analyzed quantitatively with the Rietveld program BGMN and the Profex user-interface (Doebelin and Kleeberg, 2015; Taut et al., 1998). The following mineral phases are considered for the refinements: forsterite, labradorite, bytownite, orthoclase,albite, augite, diopside, kanoite, smectite-tri, zeolite (heulandite, clinoptilolite, philippsite, dachiardite, laumontite, analcime), pyrite, ilmenite, titanomagnetite, jadeite, siderite, hematite, maghemite, anatase, titanite, schorl, chlorite, calcite, quartz, wairakite, prehnite, epidote, actinolite, garnet, grossular. Two examples of refinements are shown, in Fig. 1 for a sample containing smectite as the only clay mineral and in Fig. 2 for a sample containing smectite, chlorite, and most likely chlorite-smectite.

In samples containing a mixture of disordered clay phases (such as smectite and smectite-containing mixed layers and/or disordered 1:1 minerals), smectite content is difficult to quantify accurately by Rietveld-refinements, because of the strong correlation between parameters, including background (e.g. Raven and Self, 2017). XRD patterns of samples containing both smectite and chlorite can be fitted (Fig. 2) but the software adjusts both the chemical composition of chlorite and the relative quantity of chlorite and smectite to fit the relative intensities of the d(001) peak at 14–15 Å... and the d(002) peak at 7.2 Å.... This results in the non-uniqueness of the model parameters related to chlorite chemistry and smectite and chlorite quantities. The chemistry of chlorite can be constrained (especially the relative abundance of Fe and Mg), based on independent chemical analyses, but this requires a large number of EPMA on polished thin sections, to obtain a representative chemical composition of chlorite for the sample, which is time-consuming and expensive. Moreover, assuming that the chemistry of chlorite can be properly constrained, the combined presence of smectite and chlorite often implies the presence of a chlorite-smectite phase. This chlorite-smectite phase also contributes to the d(001) peak at 14–15 Å... and its contribution can hardly be discriminated from that of pure smectite. As a result, the fitted amount of "smectite" corresponds in reality to a quantity of smectite + chlorite-smectite. Since only the quantity of smectite layers matters for the comparison to the CEC, not the quantitiy "smectite+ chlorite-smectite", this quantification is not appropriate. Therefore, only smectite quantifications for samples containing no other clay phase are used in this study for further

Table 1
Description of the 38 samples used in this study (ID = sample name). The borehole (BH) and depth (in meters) from which the samples are extracted are indicated in Columns 2 and 3. The type of analysis for which the samples are used are indicated in Columns 4 to 6, where Opt_{CEC} corresponds to the optimization of CEC determination. The CEC (in meq/100 g) and the smectite weight per cent measured by XRD are given in Columns 7 and 8. The lithology (Litho) is given in Column 9: hyalo = hyaloclastite; v. lava = vesicular lava; d. lava = dense lava; ignimb. = ignimbrite. The presence of secondary minerals (if more than 1% as quantified by XRD quantification) is indicated in Columns 10 to 17. Sm = smectite; Chl = chlorite; Heu = heulandite (zeolite); Qtz = quartz; Clc = calcite; Pyr = pyrite; Tit = titanite; Oth. = other; Laum = laumontite (zeolite); Act = actinolite; Ep = epidote; Wai = wairakite; Ana = anatase; Jad = jadeite; Sid = siderite.

ID	BH	Depth	XRD	Opt_{CEC}	EPMA	CEC	Smec XRD	Lithology	Sm	Chl	Heu	Qtz	Clc	Pyr	Tit	Oth.
L02	KH1	39.5	x			13.5	14%	breccia	x		x	x		x		
L04	KH1	42			x	17.6	17%	breccia	x		x	x		x	x	Ana
L05	KH1	45.3	x		x	6.1	5%	ignimb.	x		x	x				
L06	KH1	48.8	x	x		7.2	13%	breccia	x		x	x	x	x		
L09	KH1	60	x	x		25.9	26%	breccia	x		x	x	x	x		
L10	KH1	66	x			15.6	18%	v. lava	x			x				
L11	KH1	68.7	x			24.9	25%	d. lava	x		x	x			x	
L12	KH1	70	x		x	2.7	4%	d. lava	x				x			
L14	KH1	74.5	x			33.2	37%	hyalo	x		x	x	x	x		
L15	KH1	79.5	x			15.0	17%	d. lava	x			x				
L16	KH1	99	x			4.5	9%	d. lava	x		x	x	x		x	
L19	KH1	120.3	x			15.1	16%	d. lava	x		x	x	x	x	x	
L21	KH1	125.5	x			19.2	18%	d. lava	x			x			x	
L22	KH1	131.1	x			21.2	17%	v. lava	x		x	x	x	x	x	
L26	KH1	157.9	x		x	12.8	16%	v. lava	x			x				
L28	KH1	167.1	x		x	13.0	18%	d. lava	x		x	x		x	x	
L29	KH1	174.3	x		x	9.4	12%	v. lava	x			x	x		x	
L30	KH1	185.1	x		x	7.2	7%	dyke	x			x			x	
L31	KH1	188.5	x		x	20.0	21%	hyalo	x			x	x	x	x	
L119	KH3	60.4	x	x		45.7	49%	hyalo	x		x	x	x	x		
L40	KH5	190			x	10.9	-	v. lava	x	x		x	x			Laum
L58	KH5	429			x	3.5	-	dyke	x	x		x	x	x	x	Act
L126	KH6	295	x			53.4	67%	hyalo	x				x			Sid, Jad, Ana
L112	KH6	377.7		x		6.2	-	dyke	x	x		x	x			
L113	KH6	387		x		3.5	-	dyke	x	x		x	x			
L114	KH6	394.2		x		2.6	-	dyke	x	x		x	x			
L81	KH6	461.32	x		x	16.0	18%	d. lava	x			x				Horn
L82	KH6	486.5			x	5.8	-	d. lava	x	x			x	x	x	
L87	KH6	508.5			x	5.6	-	d. lava	x			x	x	x	x	
L91	KH6	537.2			x	8.4	-	hyalo	x			x	x	x		
L93	KH6	555			x	1.9	-	dyke	x	x		x			x	Ep
L99	KH6	587		x		34.0	-	hyalo	x	x	x					Wai, Act
L100	KH6	597.5				0.8	-	hyalo		x		x	x		x	Wai, Act
L80	KH6	448		x		4.8	-	d. lava	x	x		x				
L86	KH6	501		x		6.3	-	d. lava	x	x		x				
L89	KH6	515		x		5.0	-	d. lava	x	x		x				
L95	KH6	560	x			39.6	38%	hyalo	x				x		x	Wai, Act
L149	KH6	680	x			19.5	24%	breccia	x			x	x			

comparison to CEC determination. We consider that samples contain smectite as the only clay mineral when no 14 Åor 10 Åpeak is clearly visible. In these cases, the presence of interstratified minerals with chlorite or illite in small quantity cannot be completely ruled out but it would have a negligible effect on the total smectite quantification.

Special studies are necessary to determine the exact type of smectite present in each sample, especially regarding the type of interlayer cation (Ca^{2+} or Na^{+}), the number of water layers surrounding interlayer cations and the tri- or dioctahedral character. The tri- or dioctahedral character mainly influences the d(060) diffraction peaks, at large angles, which have a lower intensity and overlap with other minerals such as quartz. These high-angle peaks only negligibly affect (if at all) the refinement. On the other hand, small-angle peaks, especially the d(001) around 14–15 Å..., which is influenced by the composition of the interlayer space both in terms of cation and water layer, significantly affect the refinement. Therefore, the patterns are fitted with two different types of smectite phases: a phase corresponding to a "tri-octahedral smectite with interfoliar spaces filled with Ca, accompanied by two water layers", and another tri-octahedral smectite "saponite" with a more flexible interlayer distance (see detailed structure files in Appendix B). An uncertainty regarding the smectite content is calculated based on the discrepancy between the two fits. For the other minerals, the uncertainty is calculated based on the variance of the model parameters given by the software.

2.3. CEC determination

The CEC of altered volcanic rocks is determined by back-titration of the Copper-triethylenetetramine (Cu-trien) molecule, as in the original protocol designed by Meier and Kahr (1999) to measure the CEC of pure clay samples. This molecule is also used, for example, by Kaufhold and Dohrmann (2003) to determine the CEC of bentonites.

First, the sample is weighed in a beaker and then 50 ml of deionized water, measured with a volumetric flask, are added into the beaker. The few remaining water drops in the volumetric flask, after adding water to the beaker, represent an average of 0.6 ± 0.2. In order to reduce the uncertainty in the water volume, we measure the exact mass of water added to the beaker and adjust it by weight as close as possible to 50.00 g. Without this weight adjustment and measurement, the volume of water added is 49.4 ± 0.2 ml. The beaker containing the water-rock

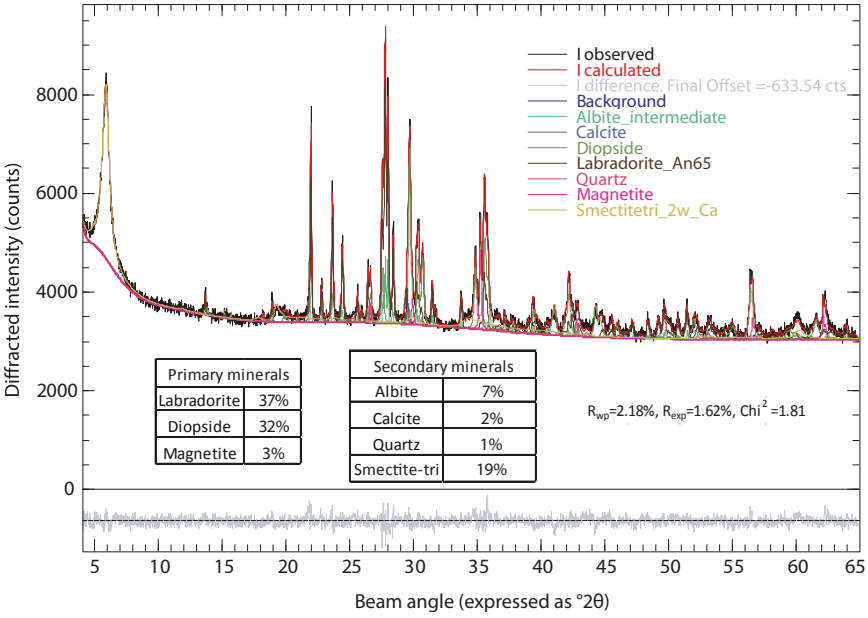

Fig. 1. XRD quantitative analysis, using the BGMN software, for sample L15 containing smectite as the only clay mineral. The resulting mineral percentages and fit quality are indicated on the graph.

mixture is then left in an ultrasonic bath for 5 minutes. Next, 10 ml of Cu-trien at about 1×10^{-2} mol/l are added with a 5 ml pipette (two steps). Details about the preparation and characterization of the exchange Cu-trien solution at 1×10^{-2} mol/l are given in Appendix C. The exchange is considered complete after 5 minutes of magnetic stirring. Since most of the exchange is expected to take place within interlayer spaces of swelling clay minerals (smectite), 5 minutes are considered sufficient. A test, carried out on sample L126, indicates that the difference in exchange yield after 5 and 60 minutes is within the instrument uncertainty, and thus not significant (see more details in Appendix C). If the CEC of higher charge clay minerals (e.g. vermiculite) were investigated, longer contact times might be needed (Von Reichenbach, 1968).

After the exchange reaction is completed, solid and liquid are separated by centrifugation. Finally, the absorbance of the supernatant solution is measured by a spectrophotometer at 578 nm. The absorbance of the Cu-trien solution before exchange, prepared independently by mixing 50 ml of deionized water and 10 ml of Cu-trien at 1×10^{-2} mol/l, is also measured. The CEC (in meq/100 g) is then calculated using Eq. (1).

$$\mathrm{CEC}_{\mathrm{lab}} = \frac{2(C_i - C_f)V}{m} \tag{1}$$

where V is the total volume of the solution (60 ml), m is the rock mass in mg and C_i and C_f are the Cu-trien concentrations in the initial and final solutions, respectively, in mol/l. C_i and C_f are calculated based on the absorbance measurements and the calibration curve presented in Appendix C.

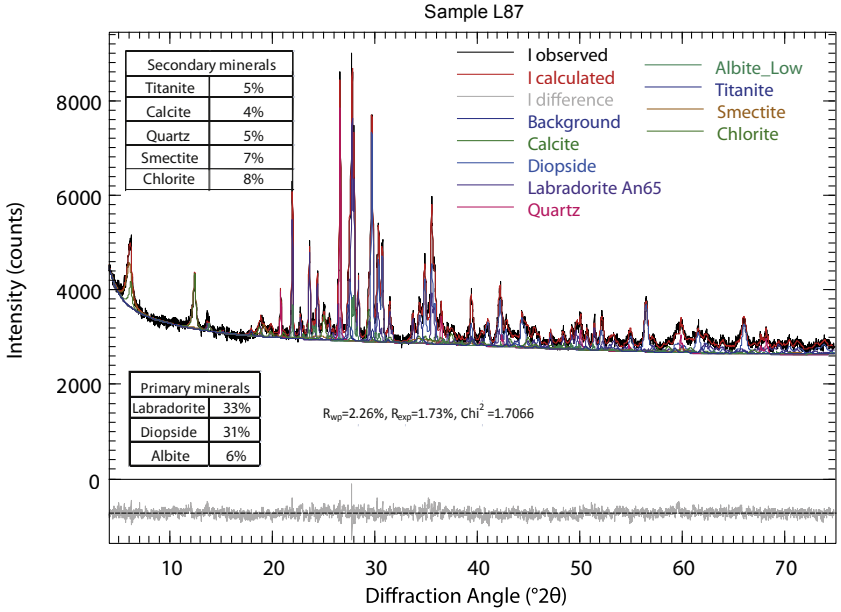

Fig. 2. XRD quantitative analysis, using the BGMN software, for a sample containing smectite, chlorite and mixed layer chlorite-smectite.

The rock mass suggested by Meier and Kahr (1999) is 200 mg but is a key parameter to be adjusted: m varies between 100 and 1000 mg in our experiments (see e.g. Dohrmann, 2006a, 2006b; Dohrmann and Kaufhold, 2009). If the CEC can be approximately estimated, e.g. thanks to XRD measurements and rapid evaluation of the d(001) peak of smectite at 14-15 Å..., the mass of rock is chosen accordingly. If the d (001) is intense compared to other peaks and no d(002) at 7.2 Å... is observed, meaning that there is little to no chlorite or chlorite-smectite, then 200 mg of rock sample are considered appropriate. If chlorite or chlorite-smectite are present in a significant amount (e.g. if d(002) is more intense than d(001)), 400 mg are used. If the d(001) peak is absent or small compared to the background, then 1000 mg are used. If no assumption on the smectite amount can be made priori to CEC determination, then a first determination with 200 mg is carried out. After the first CEC determination, the mass of rock is adjusted accordingly to the result for the next determination. At least two determinations with the same rock mass are carried out for each sample.

The protocol and equation presented here-above uses rock masses as dried at room temperature. The water content is determined independently by weighing a given mass of sample at room temperature and after drying in an oven at 105° C. Water content determinations, as well as corresponding corrections of CEC values, are presented in Appendix D.

2.4. Analysis of exchangeable cations in smectite

Chemical analysis of clay minerals were carried out on 15 polished thin sections at Géosciences Montpellier, using a CAMECA SX100 electron microprobe (22 keV, 10 nA). The 15 samples used for this analysis are indicated in Table 1.

Inductively Coupled Plasma Atomic Emission Spectroscopy (ICP-AES) analysis was carried out at ÍSOR to characterize the main cations exchanged after reaction with Cu-trien. Magnesium, calcium and sodium concentrations were determined at the wavelengths 279.079, 373.690 and 589.592 nm, respectively. The exchangeable cations of four samples are investigated by this method: L119, L96, 31 and L99. Solutions are analysed at three steps of the reaction for each sample: after mixing rock and water, after ultra-sonic bath and after exchange with Cu-trien. This allows excluding cations coming from basic water-rock interaction (e.g. dissolution of glass or minerals).

3. Results and discussion

3.1. Estimation of the laboratory uncertainty

We calculate the total uncertainty in the CEC determinations, u_{tot}, by taking into account measurement dispersion, u_{disp}, and instrument resolution, u_{instr}. The general formula, based on the rule of error propagation by Taylor expansions, is presented in Eq. (2) (e.g. Joint Committee for Guides in Metrology (JCGM)., 2008, Ku, 1966)

$$\begin{aligned} u_{\mathrm{tot}}(\mathrm{CEC}) &= \sqrt{u_{\mathrm{instr}}(\mathrm{CEC})^2 + u_{\mathrm{disp}}(\mathrm{CEC})^2} \\ u_{\mathrm{disp}}(\mathrm{CEC}) &= \sqrt{\frac{1}{n(n-1)}\sum_{i=1}^{n}(\mathrm{CEC}_i - \mathrm{CEC}_{\mathrm{avg}})^2} \\ \frac{u_{\mathrm{instr}}(\mathrm{CEC})}{\mathrm{CEC}} &= \sqrt{\left(\frac{u(V)}{V}\right)^2 + \left(\frac{u(m)}{m}\right)^2 + \left(\frac{u(C_i - C_f)}{C_i - C_f}\right)^2} \end{aligned} \quad (2)$$

where n is the number of measurements for each sample, usually two or three, and CEC_{avg} the average of the n measurements. u(V) and u(m) are the uncertainties in volume and mass, respectively, and u($C_i - C_f$) the uncertainty in the difference between initial and final concentrations. We explain below how these three terms are calculated.

The uncertainty in rock weighing is estimated to $u(m) = 0.5$ mg, based on the variations of the last digit of the scale.

The uncertainty regarding the total volume (60 ml) is calculated by propagating the uncertainty of the three volumes measured (see Eq. (3)).

$$\begin{aligned} V &= V_{\mathrm{wat}} + 2V_{\mathrm{pip}} \\ u(V) &= \sqrt{u(V_{\mathrm{wat}})^2 + 2u(V_{\mathrm{pip}})^2} \end{aligned} \quad (3)$$

where $u(V_{wat})$ is the uncertainty regarding the 50 ml of water and $u(V_{pip})$ is the uncertainty regarding the 5 ml of Cu-trien (measured twice with a 5 ml micropipette to obtain 10 ml). $u(V_{pip})$ is estimated to 0.02 ml (for each pipetting step). $u(V_{wat})$ is reduced from 0.20 to 0.01 ml when the exact mass of water added to the rock is measured. This results in a total volume uncertainty $u(V) = 0.03$ ml, while it amounts to $u(V) = 0.20$ ml when the water is direcly added from the volumetric flask without further verification.

The calculation of the uncertainty in the concentration difference is presented in Eq. (4).

$$\begin{aligned} A_i - A_f &= L(C_i - C_f) \\ \left(\frac{u(C_i - C_f)}{C_i - C_f}\right)^2 &= \left(\frac{u(A_i - A_f)}{A_i - A_f}\right)^2 + \left(\frac{u(L)}{L}\right)^2 \\ &= \frac{u(A_i)^2 + u(A_f)^2}{(A_i - A_f)^2} + \left(\frac{u(L)}{L}\right)^2 \end{aligned} \quad (4)$$

where A_i and A_f are the initial and final absorbance, respectively, and $u(A_i) = u(A_f)$ their respective uncertainty. L is the slope of the calibration curve (absorbance versus concentration, see Appendix C) and $u(L)$ is calculated using the error propagation rule presented above, with an expression of L based on the two extreme calibration points and their respective uncertainty (Joint Committee for Guides in Metrology (JCGM)., 2008).

The uncertainty in the absorbance measured by the spectrophotometer is considered to be 0.001, based on the dispersion observed for repeated measurements of the same solution, as well as the spectrophotometer resolution provided by the manufacturer.

The total instrument relative uncertainty is given by Eq. (5).

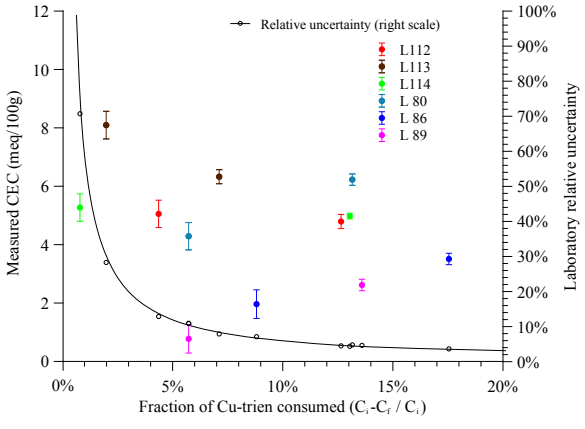

Fig. 3. CEC determined and relative laboratory (instrument) uncertainty as a function of the fraction of Cu-trien consumed. The circles with error bars correspond to the CEC determined (left axis) and the stars to the uncertainty, as calculated in Eq. (5) (right axis). An increased fraction corresponds to an increased mass of rock initially present. Each color corresponds to one sample. The six samples used in this figure have a CEC lower than 5 meq/100 g. Error bars (both positive and negative) are calculated as the product of the CEC determined by the relative uncertainty.

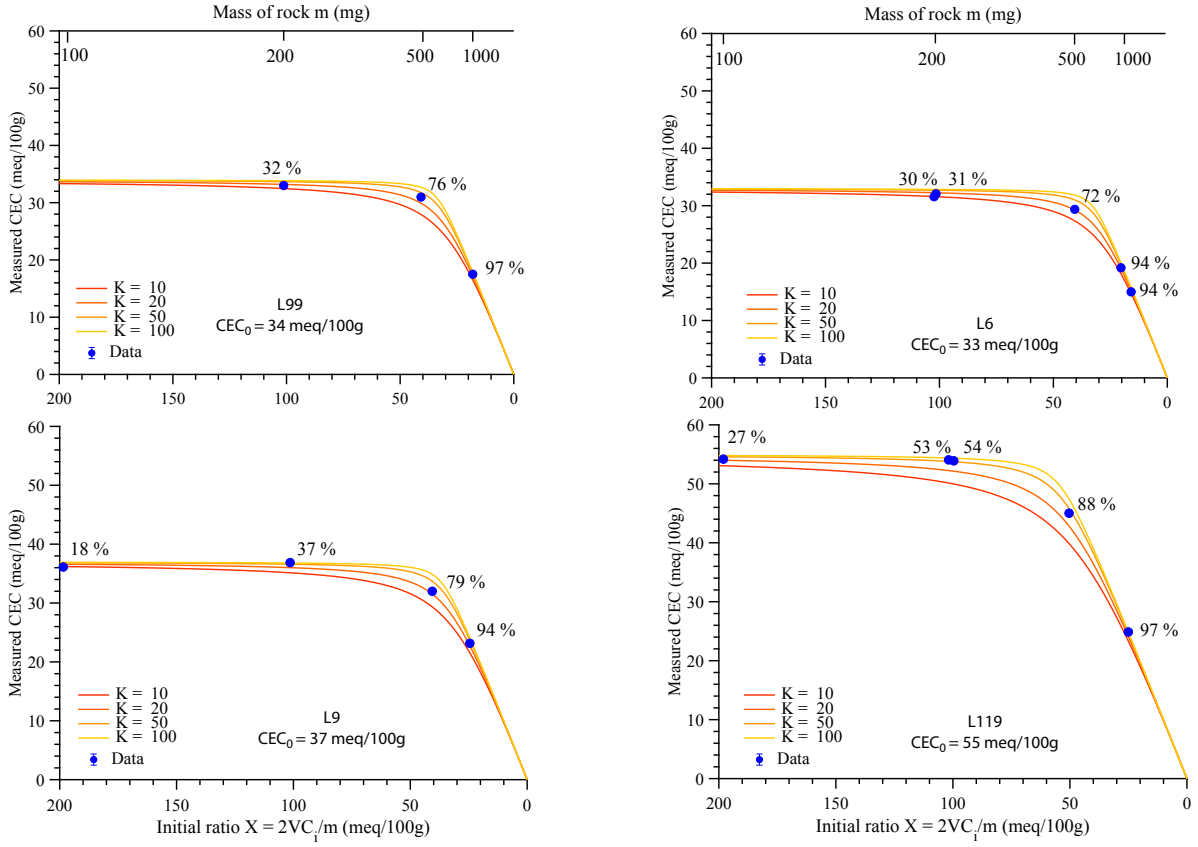

Fig. 4. CEC determined with different initial ratios of reactants (Cu-trien and rocks) for four samples (blue filled circles): L99, L9, L6 and L119. The initial concentration of Cu-trien is about constant ($(1.6 \pm 0.1) \times 10^{-3}$ mol/l) but the mass of rock varies between 100 and 1000 mg (top x-axes). The fraction of Cu-trien consumed at the end of the reaction is marked aside each determination. The laboratory (instrument) uncertainty is smaller than the symbols. The model predictions for four values of the thermodynamic constant K are also displayed as plain lines with a warm colorscale. The values of K are chosen on a trial-error basis, assuming $K \geq 1$. The value of CEC_0 used for these predictions is marked together with the name of the sample, and corresponds to the CEC determined for the highest initial ratios. The values of CEC_0 may differ from the values given in Table 1 because a different set of powder is used for these determinations than for the determinations related to the correlation with XRD refinements.

$$\frac{u_{\text{instr}}(\text{CEC})}{\text{CEC}} = \sqrt{\frac{u(V_{\text{flask}})^2 + 2u(V_{\text{pip}})^2}{V^2} + \frac{u(m)^2}{m^2} + \frac{2u(A_i)^2}{(A_i - A_f)^2} + \frac{u(L)^2}{L^2}} \quad (5)$$

3.2. Increasing the rock mass decreases the laboratory uncertainty for low-CEC rocks

Based on Eq. (5), we calculate the instrument uncertainty of CEC determinations using two different initial rock masses and the same initial Cu-trien solution for six samples. We show that the relative uncertainty decreases from up to 70% to less than 5% when increasing the rock mass from the 200 mg, as recommendend for example by Meier and Kahr (1999), to 1000 mg (Fig. 3). In the present case, increasing the rock mass from 200 to 1000 mg increases the fraction of Cu-trien consumed from less than 1% to about 13% (Sample L114 in Fig. 3).

Fig. 3 shows that the instrument uncertainty exponentially decreases with the fraction of Cu-trien consumed. This is consistent with Eq. (5), where the denominator $(A_i - A_f)^2$ controls the overall value of u(CEC) because $\frac{u(V)}{V}$ and $\frac{u(m)}{m}$ are very small compared to the third term. This is particularly true for samples with CEC lower than 5 meq/100 g, i.e. with a low smectite content. This effect is due to the limited number of digits, which can be read on the spectrophotometer (only three digits, with values always ≤ 1).

3.3. Increasing the rock mass decreases the yield of the exchange reaction for high-CEC rocks

For high-CEC samples, we observe that the CEC determined can decrease by up to 50% when the rock mass is increased, for a fixed initial volume and concentration of Cu-trien (Fig. 4). Dohrmann and Kaufhold (2009) suggest that an unsuitably small solution/solid ratio may result in a reduced selectivity of the index cation (Cu-trien in their case) and thus in an incomplete exchange of interlayer cations. A similar observation is mentioned in Ciesielski et al. (1997) for CEC measured by the Co-hex index cation and attributed to a reduced yield of the exchange reaction when the fraction of Co-hex consumed exceeds a threshold.

Here, we investigate the mathematical expression of these empirical observations, in order to predict the yield of the exchange reaction between Cu-trien and the rock sample and determine an acceptable threshold of Cu-trien consumption at the end of the reaction. The yield of the exchange reaction is taken as the relative difference between the apparent CEC determined, $CEC_{app} = CEC_{lab}$, and the maximum CEC, CEC_0. We first write the theoretical relationship between CEC_{app} and CEC_0.

For simplification purposes, we consider only one bivalent exchange

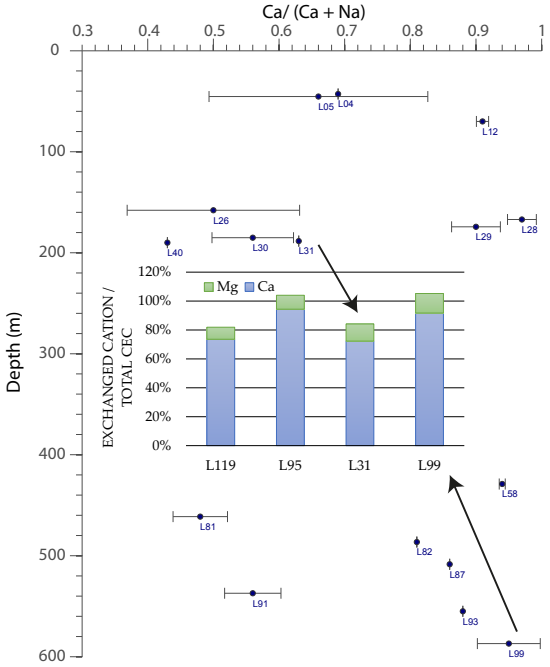

Fig. 5. Number of exchangeable Ca-ions compared, to Na-ions and Mg-ions in the interlayer spaces of smectites. The relative number of exchangeable Ca versus Na is shown as a function of sampling depth. Sixteen samples, for which chemical analysis by electron probe was carried out, are reported in this figure: L04 to L31 from KH1, L40-L58 from KH5 and L81-L99 from KH6. The error bar indicates the range of values found for the different smectite grains measured in each thin section. The histogram at the center shows the distribution of exchangeable Mg and Ca in four samples, as determined by ICP in the exchange solutions after reaction with Cu-trien. The sodium concentration in the solutions after exchange is not significantly higher than in the solutions before exchange (both before and after ultra-sonic bath) for any of the samples. Both types of determinations are available for two samples (L31 and L99).

reaction, between the Cu-trien(II) cations and Ca^{2+} cations filling the exchange sites in the rock samples. According to chemical analyses of the smectite grains carried out by EPMA, there is at least twice as much Ca^{2+} filling the interlayer spaces than Na^{+}. There is an overall abundance of Mg^{2+} in the structural formula of these saponites but EPMA does not allow differentiating Mg^{2+} in the crystal lattice from Mg^{2+} in the interlayer space. The concentrations of Mg, Ca and Na, determined by ICP in the solutions after exchange with Cu-trien for four samples, indicate that no Na has been exchanged and that exchanged-Ca and exchanged-Mg represent 72-94% and 8-13% of the total cation exchange capacity, respectively. These two types of result confirm that bivalent cations represent the majority of interlayer cations and the exchange ratio with Cu-trien(II) will be mostly 1:1 (Fig. 5).

The chemical exchange reaction, considering a majority of Ca^{2+} in the initial state, is written in Eq. (6).

$$\mathrm{Cu_{aq}} + \mathrm{Ca_{ads}} <=> \mathrm{Ca_{aq}} + \mathrm{Cu_{ads}} \tag{6}$$

The thermodynamic constant K of this exchange reaction is defined in Eq. (7).

$$K = \frac{X_{\mathrm{Cu,eq}}\, a_{\mathrm{Ca,eq}}}{X_{\mathrm{Ca,eq}}\, a_{\mathrm{Cu,eq}}} \tag{7}$$

where the subscript "eq" indicates that a chemical equilibrium is reached. $X_{Cu,eq}$ and $X_{Ca,eq}$ are the dimensionless chemical activities of cations Cu-trien(II) and Ca^{2+} filling sites in the rock at the end of the exchange reaction, respectively, and $a_{Cu,eq}$ and $a_{Ca,eq}$ are the dimensionless chemical activities of cations Cu-trien(II) and Ca^{2+} in the aqueous solution at the end of the exchange reaction, respectively. Chemical activities are defined in Eq. (8), where the subscript "eq" is removed for clarity purposes, but we assume the thermodynamic equilibrium is reached.

$$\begin{cases} @la_{\mathrm{Cu}} = [\mathrm{Cu}]_{\mathrm{aq}} \\ a_{\mathrm{Ca}} = [\mathrm{Ca}]_{\mathrm{aq}} \\ X_{\mathrm{Cu}} = \dfrac{\mathrm{meq(Cu)}}{\mathrm{mCEC_0}} \\ X_{\mathrm{Ca}} = \dfrac{\mathrm{meq(Ca)}}{\mathrm{mCEC_0}} \end{cases} \tag{8}$$

where $[Cu]_{aq}$ and $[Ca]_{aq}$ are the chemical concentrations of cations (Cu-trien(II) and Ca^{2+}, respectively) in the aqueous solution, in mol/l, and $meq(Cu)$ and $meq(Ca)$ are the quantities of cations (Cu-trien(II) and Ca^{2+}, respectively) filling sites in the rock, in millimol equivalent "meq". Each mmol of Cu-trien(II) or Ca^{2+} corresponds to 2 meq, given the double positive charge carried by each cation. a_{Cu} and a_{Ca} are normalized by an "infinite" theroretical concentration of 1 mol/l. With CEC_0 being the total CEC of the rock, in meq/g, and m the mass of rock considered in the experiment, in g, $mCEC_0$ is the total number of exchange sites available in the rock, in meq. According to these definition, $[Cu]_{aq}$, $[Ca]_{aq}$, $meq(Cu)$ and $meq(Cu)$ are given in Eq. (9).

$$\begin{cases} @l\,[\mathrm{Cu}]_{\mathrm{aq}} = C_f \\ [\mathrm{Ca}]_{\mathrm{aq}} = C_i - C_f \\ \mathrm{meq(Cu)} = \mathrm{mCEC_{app}} \\ \mathrm{meq(Ca)} = m(\mathrm{CEC_0} - \mathrm{CEC_{app}}) \end{cases} \tag{9}$$

Combining Equations (7, 8) and (9), the thermodynamic constant, K, can be written as in Eq. (10).

$$K = \frac{\mathrm{meq(Cu)}}{\mathrm{meq(Ca)}} \times \frac{C_i - C_f}{C_f} = \frac{\mathrm{CEC_{app}}}{\mathrm{CEC_0} - \mathrm{CEC_{app}}} \times \frac{C_i - C_f}{C_f} \tag{10}$$

Given that $CEC_{app} = CEC_{lab}$ and following Equations (1, 9) and (10), $meq(Cu)$ and $meq(Ca)$ can also be written, as in Eq. (11).

$$\begin{cases} @l\mathrm{meq(Cu)} = 2V(C_i - C_f) \\ \mathrm{meq(Ca)} = \dfrac{\mathrm{meq(Cu)}(C_i - C_f)}{\mathrm{KC}_f} = \dfrac{2V(C_i - C_f)^2}{\mathrm{KC}_f} \end{cases} \tag{11}$$

Assuming that all exchange sites are filled with either Cu-trien(II) or Ca^{2+} at the end of the reaction, the total CEC, CEC_0, is described in Eq. (12).

$$\begin{aligned} \mathrm{CEC_0}m &= \mathrm{meq(Cu)} + \mathrm{meq(Ca)} \\ \Leftrightarrow \mathrm{CEC_0} &= \frac{2(C_i - C_f)V}{m}\left(1 + \frac{(C_i - C_f)}{\mathrm{KC}_f}\right) \\ \Leftrightarrow \mathrm{KC}_f\mathrm{CEC_0} &= \mathrm{CEC_{app}}(\mathrm{KC}_f + C_i - C_f) \\ &= \mathrm{CEC_{app}}[(K-1)C_f + C_i] \\ \Leftrightarrow C_f &= C_i \times \frac{\mathrm{CEC_{app}}}{\mathrm{KCEC_0} + (1-K)\mathrm{CEC_{app}}} \end{aligned} \tag{12}$$

By writing C_f as a function of C_i, m, V and CEC_app (Eq. (1)), Eq. (12) gives Eq. (13).

$$\begin{aligned} C_i - \frac{\mathrm{mCEC_{app}}}{2V} &= C_i \times \frac{\mathrm{CEC_{app}}}{\mathrm{KCEC_0} + (1-K)\mathrm{CEC_{app}}} \\ \Leftrightarrow (\mathrm{KCEC_0} + (1-K)\mathrm{CEC_{app}}) &\times (2\mathrm{VC}_i - \mathrm{mCEC_{app}}) \\ &= 2\mathrm{VC}_i\mathrm{CEC_{app}} \end{aligned} \tag{13}$$

Eq. (13) can be re-written as a second-order equation of the variable $Y = CEC_{app}$ (Eq. (14)).

$$(1-K)Y^2 + (\mathrm{KCEC_0} + \frac{2\mathrm{VKC}_i}{m})Y - \frac{2\mathrm{VKC}_i\mathrm{CEC_0}}{m} = 0$$

$$\Leftrightarrow \mathrm{aY}^2 + \mathrm{bY} + c = 0 \begin{cases} @la = 1-K < 0 \\ b = \mathrm{KCEC_0} + \dfrac{2\mathrm{VKC}_i}{m} \\ c = -\dfrac{2\mathrm{VKC}_i\mathrm{CEC_0}}{m} \\ \Delta = b^2 - 4\mathrm{ac} \end{cases} \quad (14)$$

with $a < 0$ because $K > 1$ (otherwise no exchange would occur). By solving the second-order equation (14) and keeping only the positive solution (which remains positive even if $K < 1$), we obtain an explicit function for CEC_{app} (Eq. (15)).

$$Y = \mathrm{CEC_{app}} = \frac{-b+\sqrt{\Delta}}{2a} = \frac{-(\alpha+X)+\sqrt{(\alpha+X)^2+2\alpha\beta X}}{\beta} \begin{cases} @lX = 2V\dfrac{C_i}{m} \\ \alpha = \mathrm{CEC_0} \\ \beta = 2\dfrac{1-K}{K} < 0 \end{cases} \quad (15)$$

The function of X presented in Eq. (15) increases monotoneously with X and approaches asymptotically the value of CEC_0. Since X is inversely proportional to m, the sequence of equations presented here predicts that an increase of rock mass (all other things being equal) will increase the difference between CEC_{app} and its asymptote CEC_0, which is equivalent to reducing the yield of the exchange reaction.

This model based on simple assumptions predicts the observations of Dohrmann and Kaufhold (2009) and Ciesielski et al. (1997) that the sample mass is the determining factor for optimum precision of the CEC and exchangeable cations, if the same volume of solution is used. The function presented in Eq. (15) can also predict our experimental observations for samples with large smectite volumes, provided that CEC_0 is chosen accordingly (Fig. 4). Values of K in the range 10-100 are consistent with the observations but the value of K does not affect the predictions much. Given the limited sensitivity of the model to the value of K, we do not attempt here to evaluate this constant precisely. We only suggest that the range 20-50 is appropriate to describe the reaction taking place in the samples shown in Fig. 4. In reality, K depends not only on the cation exchanged but also on the type of minerals and of sites involved in the exchange (Tertre, 2014; Reinoso-Maset et al., 2012; Durrant et al., 2018; Robin et al., 2015, 2017).

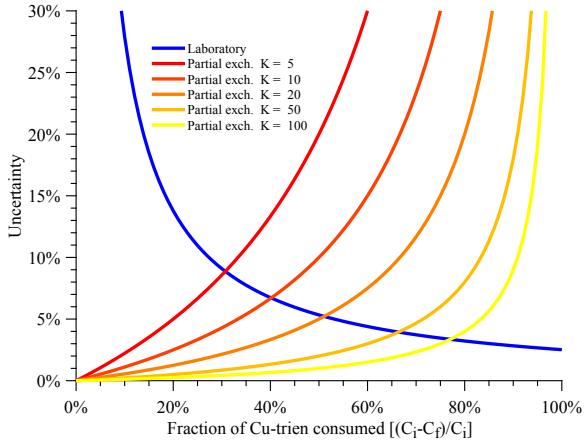

Fig. 6. Laboratory uncertainty and "partial exchange" systematic error, as a function of the fraction of Cu-trien consumed.

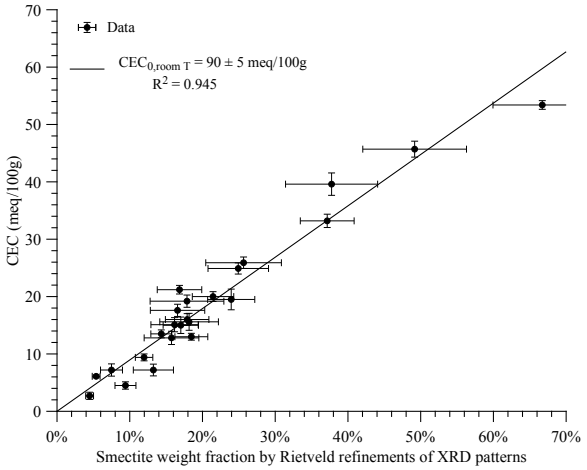

Fig. 7. CEC determinations versus smectite quantification by Rietveld refinements of XRD patterns, for samples where smectite is the only swelling clay mineral. The slope and regression coefficient of the fitting line is given in the legend. CEC determinations and XRD scans are carried out on the exact same powders. The same figure, using CEC values corrected for the water content in each sample, is presented in Appendix D.

3.4. Optimization of the CEC determination

Equations (16) and (17) describe the "partial exchange" systematic error and "instrument" uncertainty, as a function of the fraction $x = \frac{C_i - C_f}{C_i}$ of Cu-trien consumed after the reaction, based on Equations (12) and (5), respectively.

$$\mathrm{Err_{partial}} = \frac{\mathrm{CEC_0} - \mathrm{CEC_{app}}}{\mathrm{CEC_{app}}} = \frac{(C_i - C_f)}{\mathrm{KC}_f} = \frac{x}{K(1-x)} \quad (16)$$

$$\frac{u_{\mathrm{instr}}(\mathrm{CEC})}{\mathrm{CEC}} = \sqrt{\frac{u(V_{\mathrm{flask}})^2 + 2u(V_{\mathrm{pip}})^2}{V^2} + \frac{u(m)^2}{m^2} + \frac{2u(A_i)^2}{(A_i - A_f)^2} + \frac{u(L)^2}{L^2}} = \sqrt{\frac{u(V_{\mathrm{flask}})^2 + 2u(V_{\mathrm{pip}})^2}{V^2} + \frac{u(m)^2}{m^2} + \frac{2u(A_i)^2}{x^2A_i^2} + \frac{u(L)^2}{L^2}} \quad (17)$$

Depending on the value of the thermodynamic constant, K, the optimal fraction of Cu-trien consumed to minimize both the instrument uncertainty and the partial exchange systematic error is somewhere between 30% and 80% (Fig. 6), as suggested empirically for CEC determinations with Co-hex on soils by Orsini and Remy (1976). Since it is not possible to determine K with these simple measurements, we consider most resonable to aim at 30% consumption (most pessimistic value for K) because the decrease of $u_{instr}(CEC)$ beyond 30% is less important.

3.5. Quantification of smectite weight fraction in altered volcanic rocks

A linear correlation is found between the smectite weight fraction and the CEC of altered volcanic rocks, where the only swelling clay mineral is smectite (Fig. 7). Due to the complexity of Rietveld refinements in whole rock samples where several types of clays are involved, only samples where a satisfying fit is obtained using only the clay phase "Smectite tri-octahedral with interlayers filled with Ca and 2 water layers" are reported in this figure. In samples where a peak at 7–7.5 Å... is observed (typical of chlorite and chlorite-smectite), the smectite quantity derived from the fit is considered irrelevant for a quantitative

comparison with the CEC. An uncertainty in the quantification is calculated by fitting the same diffraction patterns with a different smectite phase:"Saponite with interlayers filled with undetermined cations and 2 water layers". The quantities derived from the fit with the less constrained saponite are systematically lower. The linear fit to the observations shown in Fig. 7 has a slope CEC_{smec} = 90 ± 5 meq/100 g and a regression coefficient R^2 = 0.945. The slope is consistent with the known range of CEC for smectite, 80-120 meq/100 g, caused by the permanent negative charge of the crystal lattice, in the range 0.3-0.6 per half unit cell $Si_4O_{10}(OH)_2$, that is compensated for hydrated interlayer cations (e.g. Bouchet et al., 2000). A slope of $CEC_{smec,105°C}$ = 94 ± 5 meq/100 g, with a regression coefficient of 0.952, is found when CEC values are corrected for the water loss at 105° C (Appendix D).

The contribution of other minerals to the CEC determined is also investigated, in particular zeolites, illite and chlorite. Zeolites can be divided into two groups: (i) "rigid" zeolites (e.g. laumontite, mesolite, analcime, natrolite and scolecite), whose chemical formula is well-defined and in which extra-framework (other than Al and Si) cations cannot be exchanged and only the water content can vary and (ii) "flexible" zeolites (e.g. heulandite, chabazite and clinoptilolite), which exhibit a wide and continuous range of extra-framework cation composition. Although, CEC of heulandite and clinoptilolite can reach up to 300 meq/100 g (Fridriksson et al., 2004), CEC determinations by the method developed here result in CEC values in the range 0.5-1.5 meq/100 g (Lévy et al., 2018). Therefore, the contribution of zeolites to the CEC determined by Cu-trien in altered volcanic samples is negligible.

The CEC of pure illite [Beavers Bend illite - Mankin and Dodd (1961)] was also determined using the same method (back-titration by Cu-trien), yielding 4 meq/100 g, which confirms the quasi-absence of non-mica layers in the illite sample (Mankin and Dodd, 1961) and the negligible CEC of pure illite compared to pure smectite (Hower and Mowatt, 1966). Moreover, the CEC of samples containing large amounts of chlorite, as well as in some cases wairakite and other "high-temperature" alteration minerals (epidote, actinolite), but no hint of smectite, is always lower than 0.5 meq/100 g, when determined by this method (Lévy et al., 2018). We conclude that the linear trend presented in Fig. 7 can be used to estimate the weight fraction of smectite in altered volcanic samples containing a wide range of minerals. This weight fraction also includes smectite layers in a mixed-layer chlorite-smectite or illite-smectite.

As mentioned in Section 2, the grain size of rock powders used for determinations presented in Fig. 7 is strictly below 250 μm. This is a requirement to transform the CEC determination into absolute smectite weight fraction. The presence of larger grains (e.g. millimetric size) means less surface exposed during the exchange reaction, which might result in reduced smectite accessibility by the exchange solution and thus smaller CEC determined (Kaufhold et al., 2012). We do observe a discrepancy of 20% for a high-CEC sample, L99, between the CEC determined at the same initial and optimal conditions but ground to two different sizes. These two determinations were carried out by the Institut National de Recherche en Agronomie in Arras (France), which performs accredited determinations of CEC on soil: (i) using their standard size ($\leq$ 2 mm) and (ii) using a smaller size ($\leq$ 250 μm) at our request.

Finally, heterogeneity of core samples from geothermal areas may cause significantly different CEC values depending on which lateral face of the cylindrical plug is used for the powder. Therefore, crushing and mixing together as much rock sample as possible (e.g. from the two lateral faces) is recommended.

4. Conclusions

In this study, we suggest a modified protocol to minimize the uncertainty in CEC determinations with the Cu-trien method and thus to quantify the smectite content in altered volcanic rocks. We observe that using a fixed mass of sample for rocks covering a wide range of smectite content may cause a relative uncertainty of up to 70% for samples with a low smectite content. We also show that XRD on randomly oriented powders is not sufficient for smectite quantification in samples containing other disordered clay minerals (including smectite-bearing mixed-layers) and/or chlorite. We establish that the fraction of Cu-trien consumed at the end of the reaction needs to be optimized in order to minimize the total uncertainty in the CEC determination. Instrument uncertainty and systematic "partial exchange" errors are anti-correlated with varying fractions of Cu-trien consumed. We suggest that a value of 30% for this fraction is optimal, as a rule. Finally, we show a linear correlation between the CEC, determined with an adequate Cu-trien consumption, and the smectite weight fraction determined by XRD, for 24 samples containing smectite as the only swelling clay mineral. Our study provides the geothermal industry with a simple method to quantify the smectite weight fraction (pure smectite or expandable layers in mixed-layer clays) of powders from all kinds of altered volcanic rocks. Different spectrophotometric back-titration methods, using for example the Cobalti-hexamine (III) molecule, can be used in the same manner for smectite quantification, since a whole range of thermodynamic constants are considered for the cation exchange reaction.

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Data availability: the.xrdml files (X-ray diffraction patterns) used for Rietveld quantitative analysis are available as Supplementary Material.

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Appendix A. Effect of preferred orientation on X-ray diffraction patterns: the case of heulandite

The effect on XRD patterns of preferred orientation when samples are back-loaded onto sample holders, is illustrated in Fig. A.1. The relative intensity of the heulandite peak at low angle (about 10 ° 2θ), compared to the other peaks, is much higher when sample is back-loaded. The residuals at the end of refinement (gray signal under each pattern) are also higher in this case.

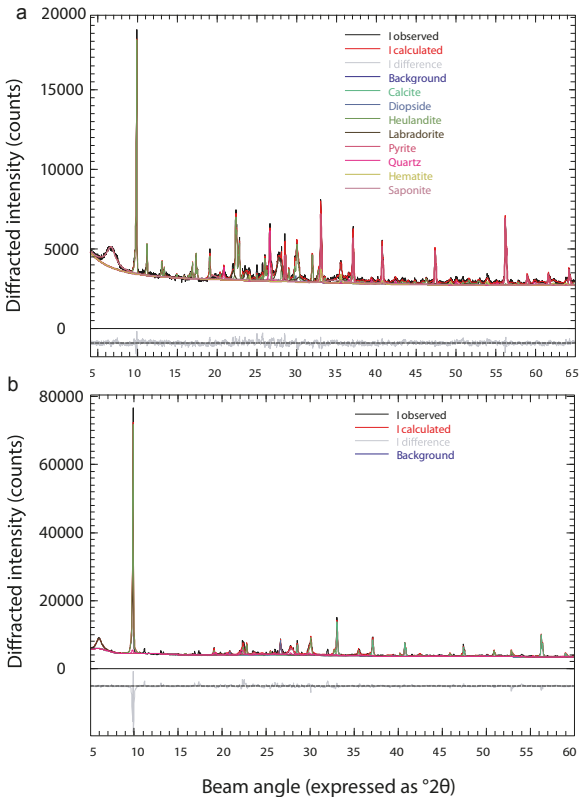

Fig. A.1. Effect of preferred orientation on X-ray diffraction pattern for a sample (L02) containing a large amount of heulandite (zeolite). The upper pannel shows the diffraction pattern when the powder is carefully grained and front-loaded. The lower pannel shows the diffraction pattern for the exact same powder but back-loaded. The diffraction peak of heulandite at angle 10 °2θ is four times higher in the lower panel and cannot be correctly fitted.

Appendix B. Structure files for smectite and chlorite in BGMN (Rietveld refinements)

B.1 Chlorite with Fe/Mg ratios constrained

```
PHASE=CHLORITE_Lea //
SpacegroupNo=12 HermannMauguin=C12/m1 //
PARAM=pa=0.6_0.5^0.8
PARAM=pb=0.6_0.5^0.8
PARAM=pc=0.6_0.5^0.8
PARAM=A=0.52558_0.51^0.55 PARAM=B=0.945_0.91^0.98
PARAM=C=1.42543_1.38^1.44
PARAM=BETA=95.587_94.5^98 //
RP=4 PARAM=B1=0_0^0.01 PARAM=k1=0_0^0.1 // PARAM=k2=0_0^0.00001 GEWICHT=SPHAR2
GOAL:chlorite2b=GEWICHT*ifthenelse(ifdef(d),exp(my*d*3/4),1)
E=(MG+2,FE+2(pa)) Wyckoff=a TDS=0.01
E=(MG+2,FE+2(pb)) Wyckoff=g y=0.6678 TDS=0.01
E=O-2 Wyckoff=i x=0.3150 z=0.9257 TDS=0.01
E=O-2 Wyckoff=j x=0.1890 y=0.1667 z=0.0774 TDS=0.01
E=(SI+4(0.6560),AL+3(0.3440)) Wyckoff=j x=0.2248 y=0.1669 z=0.1937 TDS=0.01
E=O-2 Wyckoff=i x=0.8030 z=0.7643 TDS=0.01
E=O-2 Wyckoff=j x=0.5110 y=0.2280 z=0.2363 TDS=0.01
E=O-2 Wyckoff=i x=0.8280 z=0.5711 TDS=0.01
E=O-2 Wyckoff=j x=0.1310 y=0.3463 z=0.4285 TDS=0.01
E=(MG+2,FE+2(pc)) Wyckoff=h y=0.8336 TDS=0.01
E=AL+3(0.9650) Wyckoff=d TDS=0.01
E=H Wyckoff=i x=0.2887 z=0.8552 TDS=0.02
E=H Wyckoff=i x=0.8377 z=0.6339 TDS=0.02
E=H Wyckoff=j x=0.1586 y=0.3359 z=0.3698 TDS=0.02
pMg=1- (pa+pb+pc)
GOAL=pMg
```

B.2 Saponite loosely constrained

```
PHASE=Saponite2wTest SpacegroupNo=5 HermannMauguin=C121
PARAM=A=0.53_0.525^0.535 B=A*sqrt(3) PARAM=c0=1.5_1.2^1.58
// old PARAM=c0=1.28_1.15^1.35
BETA=100.1
pi==2*acos(0)
RP=4
layer==10 // layer: factor for elongation in c direction
C=c0*layer // C: lattice parameter c for supercell
PARAM=b10=0.002_0^0.015 // isotropic broadening of hkl reflections
PARAM=b11=0.08_0^0.15 // separate broadening of 00l reflections
B1=ifthenelse(and(eq(h,0),eq(k,0)),b10+b11,b10)
// K20: strain broadening of hkl lines
PARAM=K20=0.000026_0.00001^0.0001
// K21: strain broadening of 00l lines
// changer pour ressembler nontronite de 0.001 a 0.002
PARAM=K21=0_0^0.002
breit2=1/sqr(C) // additional l-dependent broadening to avoid "ripples"
PARAM=GEWICHT=0.0_0 // refining the scale factor
// definition of the helper variable "Saponite..."
// for calculation of phase abundances
GOAL:Saponite2wTest=GEWICHT
// squared lorentzian (Gauss-like) broadening
B2=cat(R2==sqr(h/A)+sqr(k/B),Z2==max(sqr(sk)-R2,0),
orientierung2==Z2/sqr(sk),
ifthenelse(and(eq(h,0),eq(k,0)),K21*sqr(sk),K20*sqr(sk)+breit2*orientierung2))
//
// scaling of classes (00l und hkl) and removal of redundant 00l reflections
GEWICHT[1]=GEWICHT*ifthenelse(and(eq(h,0),eq(k,0)),
ifthenelse(mod(l,layer),0,layer),1)
//
// === occupancies ======================
//
// – interlayer ———
PARAM=pINT=0.3_0.2^0.4
pOZ=pINT
//
// ==== rigid body of the interlayer complex ======================
// cation, squared surrounded by 4 oxygen (water)
// definition of the positions in cartesian co-ordinates
//
dCAO=0.241 // distance cation - oxygen
//
set(ECA,0,0,0) // cation in the middle of the interlayer
// ajout de EOZ1 et EOZ2 comme dans nontronite15
set(EOZ1,0,0,dCAO) // O above
set(EOZ2,0,0,-dCAO) // O below
set(EOZ3,dCAO,0,0)
set(EOZ4,-dCAO,0,0)
set(EOZ5,0,dCAO,0)
set(EOZ6,0,-dCAO,0)
xx=0.69 // shifting parameter of the interlayer complex in x, fixed
yy=0.21 // shifting parameter of the interlayer complex in y, fixed
fi1=0 // the 3 Eulerian angles for rotation of the interlayer complex, fixed
fi2=0
fi3=-18
T(xx,yy,0.5*c0*sin(pi*BETA/180),fi1,fi2,fi3,ECA,EOZ1,EOZ2,EOZ3,EOZ4,EOZ5,EOZ6)
// shifting and rotation of the rigid body
//
// – isotropic temperature factors (nm^2), estimated ——————-
//
tdsint=0.01
```

```
tdsH2O=0.02
tdsoct=0.005
tdstet=0.003
tdso=0.007
//
// — positions ——————————
// trioctahedral coordinates from phlogopite ICSD 6259
// absolute positions in c-direction [nm]
// to avoid a stretching/shortening of the TOT layer by varying c0
//
zT=0.2708
zO11=0.112
zO12=0.104
zO2=0.328
//
E=MG+2 Wyckoff=a y=0.0 TDS=tdsoct // trans
E=MG+2 Wyckoff=a y=0.6673 TDS=tdsoct // cis
E=MG+2 Wyckoff=a y=0.3327 TDS=tdsoct // cis
E=(SI+4(0.93),AL+3(0.07)) Wyckoff=c x=0.9238 y=0.8335 z=zT/(layer*c0) TDS=tdstet
E=(SI+4(0.93),AL+3(0.07)) Wyckoff=c x=0.9238 y=0.1665 z=zT/(layer*c0) TDS=tdstet
E=O-1 Wyckoff=c x=0.979 y=0.0 z=zO2/(layer*c0) TDS=tdso
E=O-1 Wyckoff=c x=0.671 y=0.2315 z=zO2/(layer*c0) TDS=tdso
E=O-1 Wyckoff=c x=0.871 y=0.1668 z=zO11/(layer*c0) TDS=tdso
E=O-1 Wyckoff=c x=0.871 y=0.8332 z=zO11/(layer*c0) TDS=tdso
E=O-1 Wyckoff=c x=0.363 y=0.0 z=zO12/(layer*c0) TDS=tdso
E=O-1 Wyckoff=c x=0.6710 y=0.7685 z=zO2/(layer*c0) TDS=tdso
//
// list of interlayer positions
// change NA to CA et 1(pINT) to 2(pINT) + ajout de 2 ligne EOZ1 et EOZ2
E=CA+2(pINT) Wyckoff=c x=X(ECA) y=Y(ECA) z=Z(ECA) TDS=tdsint
E=O-2(pOZ) Wyckoff=c x=X(EOZ1) y=Y(EOZ1) z=Z(EOZ1) TDS=tdsH2O
E=O-2(pOZ) Wyckoff=c x=X(EOZ2) y=Y(EOZ2) z=Z(EOZ2) TDS=tdsH2O
E=O-2(pOZ) Wyckoff=c x=X(EOZ3) y=Y(EOZ3) z=Z(EOZ3) TDS=tdsH2O
E=O-2(pOZ) Wyckoff=c x=X(EOZ4) y=Y(EOZ4) z=Z(EOZ4) TDS=tdsH2O
E=O-2(pOZ) Wyckoff=c x=X(EOZ5) y=Y(EOZ5) z=Z(EOZ5) TDS=tdsH2O
E=O-2(pOZ) Wyckoff=c x=X(EOZ6) y=Y(EOZ6) z=Z(EOZ6) TDS=tdsH2O
```

B.3 Tri-octahedral smectite, interlayer spaces filled with Ca and two water layers

```
PHASE=Smectitetri_2w_Ca
SpacegroupNo=5 HermannMauguin=C121
PARAM=B=0.93_0.900^0.930 A=B/sqrt(3)-0.0015 PARAM=c0=1.50_1.42^1.6
BETA=100.2
pi==2*acos(0)
RP=4
layer==10 // layer: factor for elongation in c direction
C=c0*layer // C: lattice parameter c for supercell
PARAM=b10=0.002_0^0.015 // isotropic broadening of hkl reflections
PARAM=b11=0.03_0^0.1 // separate broadening of 00l reflections
B1=ifthenelse(and(eq(h,0),eq(k,0)),b10+b11,b10)
// K20: strain broadening of hkl lines
PARAM=K20=0.000026_0.00001^0.0001
// K21: strain broadening of 00l lines
PARAM=K21=0_0^0.001
breit2=1/sqr(C) // additional l-dependent broadening to avoid "ripples"
PARAM=GEWICHT=0_0 // refining the scale factor
// definition of the helper variable "smectite..."
// for calculation of phase abundances
GOAL:Smectitetri2wCa=GEWICHT*ifthenelse(ifdef(d),exp(my*d*3/4),1) //
// squared lorentzian (Gauss-like) broadening
B2=cat(R2==sqr(h/A)+sqr(k/B),Z2==max(sqr(sk)-R2,0),orientierung2==Z2/sqr(sk),
ifthenelse(and(eq(h,0),eq(k,0)),K21*sqr(sk),K20*sqr(sk)+breit2*orientierung2))
//
```

```
// scaling of classes (00l und hkl) and removal of redundant 00l reflections
GEWICHT[1]=GEWICHT*ifthenelse(and(eq(h,0),eq(k,0)),.....ifthenelse(mod(l,layer),0,layer),1)
//
// === occupancies =====================
// — octahedra position ——-
pMG=0.15_0.1^0.3 PARAM=pFE=0.06_0^0.3 pAL=(1-pMG-pFE)
PARAM=ptrans=1.0_0.0^1.0
// mixing parameter for cis- and trans-vacancy; 0 = > trans-vacant
//
// — interlayer ——————
PARAM=pCA=0.15_0.1^0.3
pOZ=pCA
//
// ==== rigid body of the interlayer complex =====================
// cation, octahedrally surrounded by 6 oxygen (water)
// definition of the positions in cartesian co-ordinates
//
dCAO=0.241 // distance cation - oxygen
//
set(ECA,0,0,0) // cation in the middle of the interlayer
set(EOZ1,0,0,dCAO) // O above
set(EOZ2,0,0,-dCAO) // O below
set(EOZ3,dCAO,0,0)
set(EOZ4,-dCAO,0,0)
set(EOZ5,0,dCAO,0)
set(EOZ6,0,-dCAO,0)
xx=0.7 // shifting parameter of the interlayer complex in x
yy=0.2 // shifting parameter of the interlayer complex in y
// the first two Eulerian angles for rotation of the interlayer complex,
// fixed
fi1=45
fi2=180*acos(1/sqrt(3))/pi
// fi3 (3th eulerian angle) is a rotation around the cartesian z-axis
// which is perpendicular to the xy-plane
fi3=-20
T(xx,yy,0.5*c0*sin(pi*BETA/180),fi1,fi2,fi3,ECA,EOZ1,EOZ2,EOZ3,EOZ4,EOZ5,EOZ6)
// shifting and rotation of the rigid body
//
// — isotropic temperature factors (nm^2), estimated ——————————-
//
tdsint=0.015
tdsH2O=0.025
tdsoct=0.01
tdstet=0.01
tdso=0.015
//
// — positions —————————-
// absolute positions in c-direction [nm]
// to avoid a stretching/shortening of the TOT layer by varying c0
//
zT=0.271350
zO11=0.10955
zO12=0.10553
zO2=0.33668
//
E=MG+2 Wyckoff=a y=0.0 TDS=tdsoct
E=MG+2 Wyckoff=a y=0.6673 TDS=tdsoct
E=MG+2 Wyckoff=a y=0.3327 TDS=tdsoct
E=(SI+4(0.93),AL+3(0.07)) Wyckoff=c x=0.9238 y=0.8335 z=zT/C TDS=tdstet
E=(SI+4(0.93),AL+3(0.07)) Wyckoff=c x=0.9238 y=0.1665 z=zT/C TDS=tdstet
E=O-1 Wyckoff=c x=0.979 y=0.0 z=zO2/C TDS=tdso
E=O-1 Wyckoff=c x=0.671 y=0.2315 z=zO2/C TDS=tdso
E=O-1 Wyckoff=c x=0.871 y=0.1668 z=zO11/C TDS=tdso
E=O-1 Wyckoff=c x=0.871 y=0.8332 z=zO11/C TDS=tdso
E=O-1 Wyckoff=c x=0.363 y=0.0 z=zO12/C TDS=tdso
E=O-1 Wyckoff=c x=0.6710 y=0.7685 z=zO2/C TDS=tdso
```

```
//
// list of interlayer positions
E=CA+2(pCA) Wyckoff=c x=X(ECA) y=Y(ECA) z=Z(ECA) TDS=tdsint
E=O-2(pOZ) Wyckoff=c x=X(EOZ1) y=Y(EOZ1) z=Z(EOZ1) TDS=tdsH2O
E=O-2(pOZ) Wyckoff=c x=X(EOZ2) y=Y(EOZ2) z=Z(EOZ2) TDS=tdsH2O
E=O-2(pOZ) Wyckoff=c x=X(EOZ3) y=Y(EOZ3) z=Z(EOZ3) TDS=tdsH2O
E=O-2(pOZ) Wyckoff=c x=X(EOZ4) y=Y(EOZ4) z=Z(EOZ4) TDS=tdsH2O
E=O-2(pOZ) Wyckoff=c x=X(EOZ5) y=Y(EOZ5) z=Z(EOZ5) TDS=tdsH2O
E=O-2(pOZ) Wyckoff=c x=X(EOZ6) y=Y(EOZ6) z=Z(EOZ6) TDS=tdsH2O
```

Appendix C. More details about the CEC protocol and the sources of uncertainty

C.1 Preparation of Cu-trien solutions and calibration of the spectrophotometer

Exchange solutions are prepared by mixing copper sulphate $CuSO_4$ and the organic compound tri-ethylene-tetramine "trien", in stoechiometric proportions, in a 1 L volumetric flask. The theoretical concentration of the stock solution is calculated following Eq. (18).

$$C_{\text{stock}} = \frac{\min\left(\frac{m_{\text{CuSO4}}}{M_{\text{CuSO4}}}; \frac{m_{\text{trien}}}{M_{\text{trien}}}\right)}{V_{\text{tot}}} \tag{18}$$

where m_{CuSO_4} and m_{trien} are the masses of $CuSO_4$ and "trien", in g, M_{CuSO_4} and M_{trien} are the molar masses of $CuSO_4$ and "trien", in g/mol and V_{tot} is the total volume of the solution, in L. The masses of $CuSO_4$ and "trien" are calculated to obtain a final concentration of 0.01 M Cu-trien (e.g. $m_{\text{CuSO}_4} = 1.6114$ g (anhydrous) and $m_{trien} = 1.4941$ g).

The complex "Cu-trien" is formed by stoichiometric reaction between the two compounds, so that the quantity of Cu-trien formed (in mol) corresponds to the quantity of the compound initially present in lesser quantity, the "limiting reactant". According to Stanjek and Künkel (2016), one has to avoid using an excess of "trien" in the preparation, due to a possible complexation of trien with the interlayer cations of smectite (e.g. Ca, Mg) that would prevent a later exchange with Cu-trien. Since the "trien" compound (from Sigma-Aldrich) has a purity of ≥97 % (see also in Stanjek and Künkel, 2016), and the masses are calculated as if trien were 100% pure, $CuSO_4$ is theoretically in excess in our preparation.

Most of the exchange solutions are prepared with anhydrous $CuSO_4$. Since anhydrous $CuSO_4$ is an hygroscopic compound and might have slightly rehydrated during storage, we measure the Cu concentration of both Cu-trien and $CuSO_4$ solutions by ICP (calibrated with a standard for Copper at the wavelength 324.754 nm). We compare the concentrations determined to the theoretical solutions, calculated as if it were perfectly anhydrous (Table C.1). ICP results indicate that the theoretical Cu concentrations in the $CuSO_4$ solutions, prepared with anhydrous and pentahydrated solid $CuSO_4$, are overestimated by 2.7% and underestimated by 0.8%, respectively. The relative difference in Cu concentration of 6 Cu-trien solutions, all prepared with anhydrous $CuSO_4$, caries between -4% and 4% (Table C.1). This indicates that no systematic error in the Cu concentration (due to possible rehydration and thus increase of the molar mass) shall be taken into account in the calculations.

Four independent Cu-trien stock solutions and respective dilutions are used for calibrating the spectrophotometer. This set of solutions includes the Cu-trien solutions whose Cu concentration is determined by ICP (Table C.1). It also includes two Cu-trien solutions directly prepared with the two $CuSO_4$ solutions analyzed by ICP beforehand. In these cases, the mass of solution is weighted and a mass of "trien" corresponding to a stoechiometric ratio between Cu and trien is mixed with deionized water and added to the copper sulphate solution in a 1 L volumetric flask.

The corresponding calibration curve, presented in Fig. C.1, shows that the multiplicative factor L between absorbance and Cu-trien concentration ($A = L[Cu - trien]$) is determined with satisfying accuracy ($L = 145.4 \pm 0.9$ L/mol).

Table C.1
Comparison of theoretical (conc. theo.) and ICP-determined Cu concentration (conc. ICP) in six Cu-trien solutions and two $CuSO_4$ solutions. The first six rows correspond to Cu-trien solutions whose ICP-determined Cu concentration are used for the calibration curve in Fig. C.1. The two last rows correspond to $CuSO_4$ solutions, prepared with pentahydrated and anhydrous solids.

Solution	Conc. ICP	Conc. theo	Err
	mol/l	mol/l	
Cu-trien StdA	3.45E-04	3.59E-04	4%
Cu-trien StdB	1.04E-03	9.99E-04	−4%
Cu-trien StdD	1.63E-03	1.68E-03	3%
Cu-trien StdE	2.63E-03	2.53E-03	−4%
Cu-trien 1:6 a	1.58E-03	1.52E-03	3%
Cu-trien 1:6 b	1.50E-03	1.52E-03	−2%
CuSO4 (pentahyd.)	1.01E-02	1.00E-02	1%
CuSO4 (anhyd.)	9.73E-03	1.00E-02	−3%

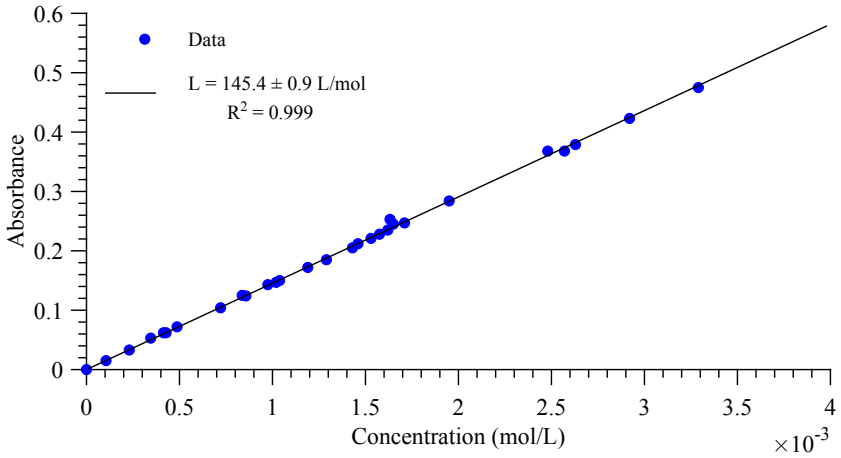

Fig. C.1. Calibration curve (absorbance versus theoretical concentration) for the Cu-trien exchange solution. The fitted line is a linear function with an intercept forced to 0. The slope *L* of the linear fit and the regression coefficient are given on the figure. The absorbance is always measured at 578 nm.

C.2 Contact time with Cu-trien

A test was carried out on sample L126 to evaluate whether longer contact times with Cu-trien might lead to increased exchange. The CEC determined after 5 and 60 min are 53.4 ± 0.7 meq/100 g and 53.1 ± 0.7 meq/100 g, respectively. We consider the difference between these two numbers not significant, which confirms that 5 min is a sufficient time for this type of sample.

Appendix D. Water content and CEC correction

CEC determinations are carried out on rock samples dried at room temperature. The CEC values presented in this study do not include a correction for the

Table D.1
Water loss at 105 °C and correction of the CEC values to take into account the water content.

ID	Water content	CEC (no correction)	CEC corrected
	wt.%	meq/100 g (room T)	meq/100 g (105 °C)
L02	3.1%	13.5	14.0
L04	2.4%	17.6	18.0
L05	1.0%	6.1	6.2
L06	2.6%	7.2	7.4
L09	6.0%	25.9	27.6
L10	3.9%	15.6	16.3
L11	2.1%	24.9	25.5
L12	1.3%	2.7	2.7
L14	3.6%	33.2	34.4
L15	1.8%	15.0	15.3
L16	1.6%	4.5	4.6
L19	2.0%	15.1	15.4
L21	2.9%	19.2	19.8
L22	2.4%	21.2	21.7
L26	0.9%	12.8	12.9
L28	2.7%	13.0	13.4
L29	1.5%	9.4	9.6
L30	1.3%	7.2	7.3
L31	2.4%	20.0	20.5
L119	7.2%	45.7	49.2
L40	0.5%	10.9	10.9
L58	0.4%	3.5	3.5
L112	1.5%	6.2	6.3
L113	1.0%	3.5	3.5
L114	0.7%	2.6	2.6
L81	1.0%	16.0	16.2
L82	0.3%	5.8	5.8
L87	0.6%	5.6	5.6
L91	0.5%	8.4	8.4
L93	0.6%	1.9	1.9
L99	4.1%	34.0	35.5
L100	4.8%	27.7	29.1
L80	1.2%	4.8	4.9
L86	1.1%	6.3	6.4
L89	1.2%	5.0	5.0
L95	5.1%	39.6	41.7
L126	7.4%	53.4	57.7
L149	3.6%	19.5	20.2

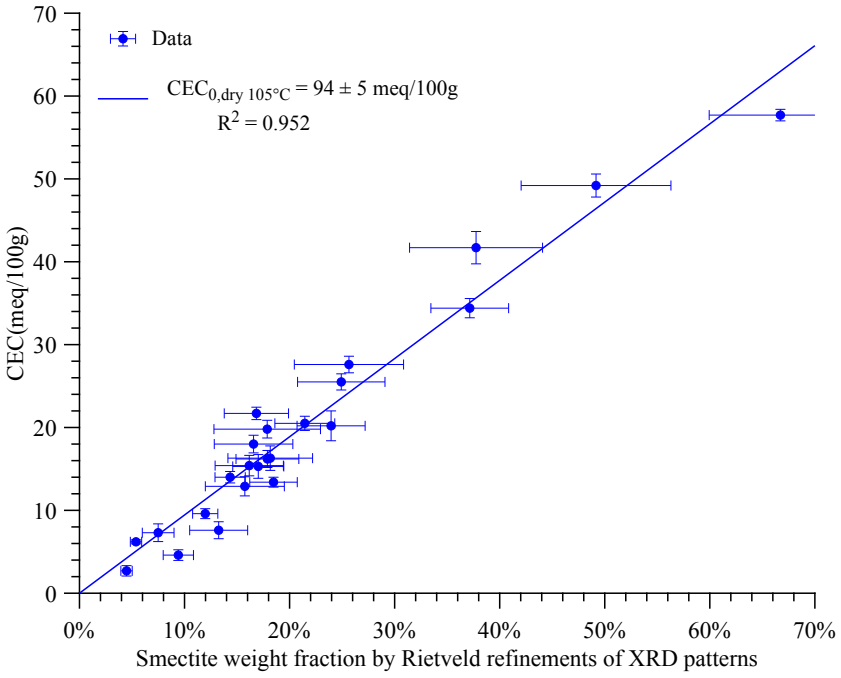

Fig. D.1. Smectite content versus CEC, after correction of the CEC value for the water content, based on the water loss at 105° C. The slope and regression coefficient are indicated on the figure.

water content in the samples. In particular, the slope presented in Fig. 7 corresponds to the average CEC of smectite in samples dried at room temperature, i.e. containing up to 7% of bound water molecules. We present in Table D.1 determinations of the water content and corrected CEC values for all samples presented in this study. The water loss is quantified by drying a given mass of each sample at 105° C. Fig. D.1 shows the correlation between CEC, as corrected for the water content, and smectite content. This slope results in a CEC of pure smectite slightly higher than when considering the uncorrected CEC values.

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